Reactivity of hydride bridges in a high-spin \([\text{Fe}_3(\mu-\text{H})_3]^{3+}\) cluster: reversible \(\text{H}_2/\text{CO}\) exchange and \(\text{Fe}^2-/\text{B}^2-\) bond metathesis\(^\dagger\)

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The tririon trihydride complex \(\text{Fe}_3\text{H}_3\text{L}(1)\) [where \(\text{L}^{3-}\) is a tris(β-diketiminate)cyclophanate] reacts with \(\text{CO}\) and with \(\text{BF}_3\cdot\text{OE}t_2\) to afford \((\text{Fe}^\text{III}\text{CO})_2\text{Fe}^{\text{II}}(\mu_3-\text{H})\text{L}(2)\) and \(\text{Fe}_3\text{F}_3\text{L}(3)\), respectively. Variable-temperature and applied-field Mössbauer spectroscopy support the assignment of two high-spin (HS) iron(iii) centers and one HS iron(i) ion in 2. Preliminary studies support a CO-induced reductive elimination of \(\text{H}_2\) from 1, rather than \(\text{CO}\) trapping a species from an equilibrium mixture. This complex reacts with \(\text{H}_2\) to regenerate 1 under a dihydrogen atmosphere, which represents a rare example of reversible \(\text{CO}/\text{H}_2\) exchange and the first to occur at high-spin metal centers, as well as the first example of a reversible multielectron redox reaction at a designed high-spin metal cluster. The formation of 3 proceeds through a previously unreported net fluoride-for-hydride substitution, and 3 is surprisingly chemically inert to \(\text{Si}-\text{H}\) bonds and points to an unexpectedly large difference between the \(\text{Fe}^2-/\text{Fe}^0\) bonds in this high-spin system.

Introduction

Beginning with Hieber’s discovery of \(\text{H}_2\text{Fe}(\text{CO})_4\) and \(\text{H}^\text{Co}(\text{CO})_4\),\(^1\) the synthesis and reactivity studies of transition metal hydrides have remained an active area of chemical research.\(^2\) Mono- and poly-nuclear metal hydrides have applications in industry (e.g., hydroformylation), electrocatalysis, and chemical synthesis.\(^2\) In addition, polynuclear transition metal hydrides have garnered substantial interest as these species have either been identified as or are proposed to be key intermediates in reactions in biological and chemical systems. For example, hydrides bound to iron centers are known or proposed in the Haber–Bosch\(^4\) and Fischer–Tropsch\(^5\) processes, dihydrogen oxidation/proton reduction by the nickel–iron and iron–iron hydrogenases,\(^6\) and the catalytic cycles for nitrogenases\(^7\) and nickel–iron carbon monoxide dehydrogenase.\(^8\) Molecular systems designed to model the multimetallic activation of \(\text{H}_2\) to generate and probe the reactivity of polynuclear metal hydrides generally use strong field donors (e.g., \(\text{CO}\) or phosphines).\(^9\) It remains unclear how parameters such as the metal coordination number and spin state tune reactivity of metal hydrides. This consideration is important in the context of heterogeneous and biological metal hydride clusters in which high- or intermediate-spin metal centers are typically present. The \(\text{NiFe}\) and \(\text{FeFe}\) hydrogenases are exceptions to the biological trend as the metal centers are low spin.

Among the most important reactions of transition metal hydrides are reductive elimination (re) and oxidative addition (oa) of \(\text{H}_2\), which are ubiquitous in organometallic chemistry, including such processes as catalytic asymmetric hydrogenation.\(^10\) The study of these reactions in polynuclear high-spin iron clusters is of particular interest because of mechanistic questions related to the reduction of \(\text{N}_2\) to \(\text{NH}_3\) by the iron molybdenum cofactor in nitrogenase from \textit{Azotobacter vinelandii}, which has recently been proposed to store reducing equivalents as hydrides that unmask low-valent iron centers upon elimination of \(\text{H}_2\) to facilitate binding and reduction of \(\text{N}_2\).\(^7\) Relatedly, the chemistry of metal hydrides is also germane to chemical hydrogen storage materials for fuel cell applications, in which repetitive cycling between \(\text{H}_2\) re and oa are important reactions for discharge and recharging, respectively.\(^11\) Ideally, the former would be stimuli-dependent (e.g., reduced pressure, light, exogenous agent) to allow for controlled and on-demand \(\text{H}_2\) release. A fundamental understanding of this type of reactivity is thus important to ongoing research efforts, even if transition metal hydrides lack sufficient energy density for use
in such applications. Exposure to N₂, reduced pressure, or visible light irradiation are known to instigate reductive elimination of two hydrido donors from transition metal complexes.¹⁰ In those cases where N₂ is involved, the interplay between N₂ binding and H₂ loss is unclear. Ideally then, compounds that can cycle between H₂ re and oa as a function of a chemical stimulus provide important insight towards tuning the energy demand for discharge and recharge of transition metal hydrides.

We previously reported the synthesis and reactivity of a series of trimetalllic trihydrido complexes, M₃(μ-H)₃L (M = Fe¹IV, Co¹IV, Zn¹II), supported by the tris(β-diketiminate)cyclophanate, L. In those reports, we highlighted the surprisingly specific reactivity of these compounds for CO₂. These metal-hydride compounds were unreactive towards other potential substrates, including nitriles, aldehydes, and ketones. Notably, the Zn¹ analog demonstrated stability even towards methanol and water. In an effort to expand the substrate scope, we report here the reactivity of Fe₂H₃L (1) with carbon monoxide, which resulted in H₂ elimination to generate a formal triiron(μ⋅μ⋅μ) compound, (Fe²CO)₂Fe³(μ₃-H)₃L, and with BF₃·OEt₂, which afforded Fe₂F₃L through a net fluoride-for-hydride substitution. To our knowledge, the H₂/CO reactivity demonstrated here is the first example of a reversible multielectron redox reaction occurring at a designed high-spin multimetallic compound. Similar reactivity under a minimal driving force is typically invoked as a hallmark of biological catalysis.

Results and discussion

Reaction of 1 with carbon monoxide (~1 atm) in THF results in a rapid color change from red to dark yellow-green (broad absorption with λmax = 320 nm, Fig. S14, ESI†) and formation of (Fe²CO)₂Fe³(μ₃-H)₃L (2) in quantitative yield by ¹H NMR spectroscopy and in good crystalline yield (Scheme 1, 69%). A new resonance corresponding with H₂ is also observed in in situ NMR spectra of reaction mixtures, implying H₂ re upon reaction of 1 with CO. This formulation of 2 is supported by single-crystal X-ray diffraction (vide infra) and combustion analysis. We also observe a strong IR absorption at 1846 cm⁻¹ in spectra of 2 synthesized from ¹²CO, which shifts to 1804 cm⁻¹ for the ¹³CO labelled 2 (2-¹³CO) (theoretical/observed ∆ν(¹²CO−¹³CO) = 41/42 cm⁻¹).

In the solid-state structure of 2, two iron centers (Fe2 and Fe3) are in a distorted trigonal pyramidal geometry (t₄ = 0.75) with each ligated by one semi-bridging CO, one ligand β-diketiminate (or nacnac) arm, and a μ₂-hydride (Fig. 1, left). The calculated values for α (0.35 and 0.39) and the Fe2/Fe3–C≡O bond angles (163.7(2)° and 162.2(2)°) agree with Curtis’ correlation¹¹ for bent semi-bridging CO ligands. However, bent Fe¹-C≡O bonds are known for terminal carbonyls (e.g., ~173° for a (nacnac)Fe²(OC)₂ complex),¹² and the geometric constraints imposed by L⁻⁻ and μ₂-H⁻ likely limit the distance and angular relationships as compared with pseudo-octahedral metal centers. The Fe2/Fe3–N₃ bond distances for 2 are significantly longer than for monometallic nacnac Fe(i) carbonyls (2.05 Å vs. 1.97–1.98 Å),¹³ suggesting differing iron spin states or steric effects. The Fe–C bond distances (1.829(2) and 1.830(2) Å) are within the range of those reported for the (β-diketiminate)iron(0) di- and tri(carbonyl) complexes (1.79–1.87 Å)¹⁵ and for [PtH₂ffFe¹(OC) (1.789(3) Å)¹⁶ but substantially longer than those reported for Peters’ trigonal bipyramidal Fe(i)–CO series (1.679–1.769 Å).¹⁷ The energy of ν(CO) for 2 is also lower than those reported for terminal Fe¹–CO species consistent with an interaction with Fe1.

Comparison of the Fe–C and C–O bond distances in 2 with those reported for other complexes featuring the monocarbonyl iron motif demonstrates that 2 is atypical with unusually long C–O bonds given the Fe–C bond lengths (Fig. S16–S17†). We cannot discount libration effects as one reason for the unexpected differences in the bond lengths, although these values are reasonable given the IR data. The (nacnac)Fe(H)(CO) units in 2 represent the second example of a four coordinate Fe¹ bearing only one CO donor, the first being [PtH₂ffFe¹(OC)].¹⁶ The Fe(H)(CO) fragment is known for P- and N-donor compounds,¹⁸ hydrogenase model compounds,¹⁹ and in metal carbonyl cluster chemistry,²⁰ but is unknown for lower than 5-coordinate iron centers. Fe1 is held within a distorted trigonal bipyramidal coordination environment comprising the μ₂-hydride and two N-atoms from the nacnac arm, and two long axial interactions of 2.481(2) Å and 2.545(2) Å to the semi-bridging CO donors.²¹ The bond metrics suggest minimal direct overlap between the iron centers in 2, which is evident from the formal shortness ratios for the Fe1 contacts (FSR = 1.24 and 1.31).²² Specifically, the Fe2–Fe3 (3.8612(4) Å) and Fe2/Fe3–Fe1 (2.9038(5), 3.0428(5) Å) distances are longer than typically associated with iron–iron bonds. The cyclophanate is distorted...
based on the bond metrics and angles relative to those of the M₅X₆L series (M = Mn, Fe, Co, Zn, Cu; X = Br⁻, Cl⁻, and H⁺), which highlights the surprising flexibility of L⁺. The (µ₄-H) Fe₅ fragment (n > 2) was previously unknown in the high-spin state (vide infra), and 2 is the first example of an internal µ₃-hydride in polyiron chemistry; face-capping µ₃-hydrides are well-known in iron carbonyl clusters.

Reactive elimination of H₂ from 1 should afford a triron complex with 20 d electrons with formal oxidation states of Fe²⁺Fe³⁻. This assignment is supported by the zero-field Mössbauer spectrum (Fig. 2), which comprises two different quadrupole doublets with a 2 : 1 area ratio. From the Mössbauer parameters, we assign the doublet with double integration as Fe¹ (δ = 0.66 mm s⁻¹; ΔE₉ = 2.60 mm s⁻¹) and corresponding to Fe2 and Fe3, whereas the parameters of the other doublet are characteristic of high-spin Fe²⁺ (δ = 0.98 mm s⁻¹; ΔE₉ = 2.15 mm s⁻¹) and attributed to Fe1. The δ values assigned to Fe2 and Fe3 are comparable to those for the high-spin (β-diketiminate)-supported iron(i) complexes with two tert-butylisocyanides or one dinitrogen bound (c.f. 0.64 or 0.62, respectively), and distinct from the analogous low-spin iron(i) di- and tri(carbonyl) and tert-butylisocyanide compounds (c.f. 0.12–0.25). This comparison suggests that 2 likely contains high-spin Fe³⁻ centers, which is rare for a CO ligated iron center with tris[pyrazolyl]borate Fe(CO)₃ compounds as the only characterized examples. Applying a magnetic field of 2 T at 4.8 K resulted in a slight line broadening for the doublet assigned to Fe(ii), revealing an internal field of less than 4 T (Fig. S19†). This is indicative of an integer-spin system with a relatively well-isolated M₅ = 0 ground state within the spin manifold. On the contrary, the doublet assigned to Fe(i) is magnetically split between −3 mm s⁻¹ and +4 mm s⁻¹, revealing a much larger internal field. Such behavior excludes the possibility of strongly coupled Fe centers. Consistently, we observe a broad signal near g = 16 in X-band EPR spectra collected in parallel mode on 2 (Fig. S13†). Similar spectra are reported for biological and other synthetic polyiron systems and are attributed to weak coupling between integer-spin multiplets. As a first approximation, we have simulated the high-field Mössbauer spectra assuming three uncoupled iron centers, two of which account for the two Fe(i) centers, with identical parameters. The overall splittings were well reproduced assuming an S = 3/2 ground state for the Fe(i). Additionally, the solution magnetic susceptibility of 2 (µ eff = 5.6 µB) is lower than that of all triiron(n) complexes supported by L⁺ (6.2–7.5 µB) and is therefore consistent with the Fe²⁺Fe³⁻ oxidation state assignment.¹³,²⁵

The formal metal oxidation state assignments and the overall cofactor charge in the resting state (E₀) of FeMoco remain unclear. Early formulations considered Mo(iv) and various ratios of Fe(ii) and Fe(iii); some of the Fe(ii) and Fe(iii) centers could be treated as mixed-valent pairs (i.e., 2Fe²⁺–Fe³⁺).¹⁹–²⁰ Currently, the Mo is proposed to be a non-Hund’s S = ½ Mo(iii) center based on high-resolution fluorescence detected Mo K-edge XAS.²⁹ Spatially resolved anomalous scattering (SpReAD) experiments support the Mo(n) assignment and suggest an [MoIIIFeIIFe²⁺₃Fe⁴⁺₃S₆Cl]⁺ cluster.²⁹ Recent DFT studies by Bjørnsson, et al. benchmarked to the Mössbauer parameters reported by Yoo, et al. are consistent with the SpReAD result or the related isoelectronic mixed-valent formalism (i.e., [Mo IIIFe²⁺²Fe⁴⁺₂Fe⁴⁺₂S₆Cl]⁻).²⁸ Similar ambiguities arise for the one-electron reduced (M²⁻) and the CO-bound forms of FeMoco, although differences in the experimental conditions (i.e., extracted FeMoco vs. holo-MoFe protein) may account for these discrepancies.²⁸,²⁹–³¹ Relatively, the extent of charge delocalization may be critical to dinitrogen activation. Local iron(i/ii) character would be required should N₂ binding and activation by FeMoco parallel synthetic systems.⁴² Our results and the possible coordination number changes for iron center(s) in FeMoco during catalytic turnover¹¹–¹³ hint that these uncertainties should be expected in the reduced states of FeMoco. For example, the isomer shift and quadrupole splittings for Fe2 and Fe3 could lead to the mistaken assignment of these centers as di-rather than mono-valent, since isomer shifts of 0.66 mm s⁻¹ are well within the range of many iron(n) species.

We previously suggested that loss of H₂ from 1 may be sterically controlled based on the large H···H distances (2.93–3.05 Å).¹⁴ H₂ re from 1 is not observed either thermally or under broad wavelength irradiation. For an H₂ re step to precede CO binding, an H₂ re/oa equilibrium would likely have to exist in which any liberated H₂ reacts rapidly to regenerate 1, but affords 2 in the presence of high concentrations of CO relative to H₂ (Scheme 2). H₂/D₂ exchange should occur in this mechanism, thereby scrambling the respective isotopic labels. No changes were observed in ¹H NMR spectra recorded on reaction mixtures of Fe₃D₃L (1-D₃) with H₂ (≈1 atm) at temperatures up to 80°C for 20 h (Fig. S5†). The lack of observable H/D exchange supports CO binding preceding H₂ loss and contrasts the exchange reported for dimeric β-diketiminate complexes. As noted in Scheme 2, we cannot exclude a tightly associated Kubas-type complex between the eliminated H₂ and the triiron species, which does not exchange with dissolved H₂ or D₂. CO can then coordinate to the Kubas-type species, liberating H₂ and generating 2. However, such tightly bound M–H₂ complexes on a high-spin metal center are rare,¹⁴ and this pathway assumed unlikely. A similar substrate-binding prior to H₂ elimination was recently reported for reaction of a di(µ-hydrido) diiron(n) complex with diazocompounds. In the simplest mechanism, retention of one hydride in 2 would arise from an

![Fig. 2 Mössbauer spectrum of 2 recorded at 80 K and zero-applied field. Black bars and colored lines represent the experimental data points and simulated quadrupole doublets, respectively (Table S1). Solid black line is a composite spectrum obtained by combining individual doublets.](image-url)
intramolecular step for H₂ loss upon reaction of 1 with CO (Scheme 2). Previously, we provided crystallographic evidence demonstrating that the β-diketiminate and benzene substituents on the ligand for two tririon complexes could interdigitate in the presence of K⁺ cations.⁴¹ A similar bis(tririon) intermediate could form during H₂ re; however, the predicted long H…H distance and the absence of K⁺ cations suggests that such an intermediate is also unlikely. Consistent with an intracenter H₂ re mechanism, we observe only H₂ and not HD in the ¹H-NMR spectra of reactions containing 1 and 1-D₃ under a CO atmosphere (Fig. S7†).

Given that we have hitherto been unable to isolate Fe¹-containing complexes for L¹⁻,⁴¹ we were intrigued if reactivity characteristic of Fe³ centers could be accessed starting from 2. To that end, we examined the conversion of 2 to 1 upon addition of H₂. Notably, we observed complete reaction of a toluene solution of 2 with H₂ at 20 bar and 80 °C to yield 1 in 75(7)% spectroscopic yield as a 6.3 : 1 mixture of 1 and an as-yet undetermined decomposition product. Although reversible H₂ re/oa both with and without coordination/dissociation of N₂ are common reactions,⁴¹¹⁴ the reversible H₂ re/oa with CO coordination/dissociation that we observe here is a very rare transformation. Despite the ubiquity of organometallic carbonyls and hydrides, this reaction has been previously reported only for one series of low-spin Group VIII HM₃(CO)₉ compounds.⁴⁶

Because the reactions of mononuclear or self-assembled dinuclear β-diketiminate iron hydrides with CO lead to low-spin polycarbonyl products whose formation is presumably irreversible,⁴⁰ we hypothesize that our use of a relatively rigid polynucleating ligand is the key factor that enables the unusual ability to convert between the two species that we observe. Additionally, the interchange between 1 and 2 represents a reversible two-electron redox reaction occurring at a high-spin polyanion cluster, and the starting species 1 demonstrates unusually high substrate specificity.¹³ As such, the reactivity demonstrated here bears resemblance to the biological activity of multmetallic cofactors; the proposed dihydride species in the E₄ state of FeMoco is stable to protonation but readily undergoes reversible re/oa of H₂. These observations suggest that the bio-inspired design of polynucleating ligands to stabilize high-spin clusters of base metals is a promising strategy for developing base-metal promoted multi-electron chemistry similar to that performed by multmetallic enzymes in biology.

If the dihydrogen reductive elimination is intramolecular, the distance between two hydride donors must necessarily decrease from the solid-state structure of 1 to form H₂. We posit that CO binds to 1 and stabilizes a structure in which two H⁻ donors are sufficiently close to liberate H₂. Fluxional coordination of hydrides⁷⁷ was proposed in 1 based on our previous reaction study, and a shift of a μ-hydride in 1 towards the internal cavity or to a μ₃ position would generate an open coordination site for CO to bind. With this in mind, we hypothesized that small Lewis acids could coordinate to one or more hydrides in 1 and stabilize such a distorted conformation of [Fe₂H₄]⁺.⁴⁸ We probed the reaction of 1 with BF₃·OEt₂ as others have isolated species with R₃B·H–M interactions.¹⁸ To our knowledge, the M–H···BF₃ adduct is rare with the only reported example for a bis(cyclopentadienyl)niobium dihydride.⁴⁹ We considered, however, the size of the R group and the need for an electron deficient borane in our choice of BF₃·OEt₂.

Reaction of 1 with BF₃·OEt₂ resulted in formation of a bright yellow solution of Fe₃([μ-F]₂)L₃, 3, (λ_max = 321 nm, 392 nm, 431 nm, Fig. S15†) in near quantitative spectroscopic yield and moderate crystalline yield (Scheme 3, 57%). Our assignment of 3 is supported by combustion analysis, HR-ESI/MS(+), and ¹H-NMR; the latter is consistent with a D₃h symmetric species in solution and on the method timescale. The solution magnetic susceptibility of 3 is 7.4 μₜ, consistent with that of similar tririon[μ] complexes.¹⁰b,⁷⁶ The solid-state structure of 3 evidences three bridging fluoride ligands and a planar [Fe₂F₂]⁺ core in 3 (Fig. 3) with comparable bond metrics to related mono- and di-metallic complexes (Table S2†) as well as those in tririon[μ] and -iron[μ] complexes of L¹⁻.¹³b,¹⁵,⁴⁶ Given the quantitative conversion of 1 to 3, the reaction is a net metathesis of Fe→H for B–F bonds. Consistently, treatment of the reaction mixture with NEt₃ allowed observation of BH₃ as the amineborane adduct, Et₃NBH₃, in ¹B-NMR spectra. To our knowledge,

| Scheme 2 | Possible associative (left) and dissociative (right) pathways for formation of 2, in which the two pathways can intersect depending on whether H₂ loss succeeds (diagonal dashed arrow) or precedes (horizontal arrow) CO coordination. |
| Scheme 3 | Preparation of Fe₃F₃L (3). |
this reaction represents the first example of a controlled B–F/Fe–H for B–H/Fe–F bond metathesis. This transformation is enthalpically disfavored (~21 kcal mol\(^{-1}\)) per B–F/Fe–H bond metathesis) based on the average empirical values for the respective bonds.\(^{39}\) The fact that this reaction proceeds at ambient temperature, however, is likely due to the ligand’s unique geometric constraints and the differences in bridging ligand properties of fluoride versus hydride. The B–F/Fe–H bond metathesis further illustrates the unexpected reactivity profile of 1 and related trihydride compounds.\(^{15}\) Previously, monometallic (nacnac)iron fluoride complexes were reported to react rapidly with trialkylsilanes to generate the corresponding iron hydride complexes.\(^{40}\) However, 3 is unreactive towards the silanes tested based on IR and \(^1\)H-NMR spectroscopy. This reaction should be only slightly enthalpically disfavored relative to the reaction of 1 with BF\(_3\), but our result is consistent with an unusually large difference between the Fe–F and Fe–H bond strengths in 3 vs. 1. Other factors such as approach of the silyl reagent, access to the transition state geometry, or poor access to the fluoride ligands may also modulate the reaction rate.

**Conclusions**

We have expanded on the prior reactivity studies reported for 1, demonstrating that CO triggers reductive elimination of H\(_2\) from 1 to generate an unusual high-spin mixed-valent triiron(III/II) complex, 2, and that reaction with BF\(_3\)-OEt\(_2\) affords the tri[µ-fluoride] species 3. These results add to the growing evidence that hydrides act as effective masks or protecting groups for low-valent high-spin metal centers. In addition, 2 reacts with H\(_2\) to regenerate 1, which is a rare example of reversible H\(_2\) re/oa involving CO and the first to occur at high-spin metal centers. Ongoing work focuses on continuing to develop the multi-electron reactivity profiles of 1 as well as that of 2.

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**Notes and references**


20 FeI is disordered in the structure solution: two positions are modelled with minor occupancies of 3.5% and 2.5% (Fig. S18, ESIF). Electron density corresponding to the μ3-hydride required partial occupancy of a light atom to model satisfactorily (96% H, 4% O); the contaminant could not be detected by combustion analysis performed on multiple samples of 2.


27 W. H. Harman, Personal communication.
